UG part I (HONS) & (Subsi)

## Assignment

Answer these questions

1. Name all the alkaline earth metals and write down their atomic no and electronic configuations.
2. Which among Ti(22), V(23), Cr(24) and Mn(25) has the maximum second ionisation energy?
3. Define electronegativity and write down factors on which it depends.
4. Define ionisation potential and explain Why
   (i) Ionisation potential of Li is greater than Na.
   (ii) Ionisation potential of Na is less than Mg.
   (iii) First ionisation energy of K is less than Ca but second ionisation energy is high.
5. Name alkali metals and give their electronic configuration.
6. Why ionisation potential of nitrogen is greater than oxygen and phosphorus both.
7. Arrange the following in decreasing order of I.P
   a) Be, B, C, N, O
8. Which has the second ionisation potential among N, C, O and F.
9. Ionisation energy is highest for which among
   a) Noble gases b) Platinum metals c) Transition elements
   d) inner transition elements.
10. Fluorine is best oxidising agent why.

## Electron affinity

Give reasons for following

i) Electron affinity value of alkaline earth metals of ~~Group~~ II is zero.

ii) Electron affinity value of nitrogen and phosphorus are lesser than carbon and silicon

(iii) The theoritical value of electron affinity of group 18 is zero.

iv) In moving from top to bottom the electron affinity value of elements decreases

v) In general electron affinity value increases on moving left to right in a period,

vi) The electron affinity ~~is less~~ of fluorine is less ~~than~~ chlorine.